

Towards Optimizing Synthesis Temperature for Microgels with Large Degree of Deswelling

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Polysaccharide microgels have been synthesized at various temperatures (Tsyn) above the LCST of the parent polymer. Microgel structure and dynamics below and above the corresponding volume phase transition has been studied with light scattering. All microgels were found to undergo a reversible 15-50-fold deswelling in volume. However, the size distribution, structure, dynamics, and deswelling ability of microgels were found to strongly depend on synthesis temperature. In this work, the attempt was made to optimize the synthesis temperature to yield more monodisperse microgels with a larger degree of deswelling and to understand the role of the Tsyn on the density of microgels.

□ Due to their environmental responsivity and low toxicity, our microgels have potential to be used in **biomedical applications** including targeted drug delivery

Abstract

Polysaccharide Microgels

□ **Microgels** are gel clusters formed by the chemical crosslinking of polymer chains

 We synthesized microgels with **hydroxypropylcellulose (HPC), an FDA-approved amphiphilic polysaccharide**

Using SLS and DLS to explore the effect of T_{syn} on:

 Since microgels **inherit the properties of their parent polymer**, our particles exhibit a volume phase transition at high temperatures

> • ΔT appears to strongly impact both the swollen size and relative loading capacity of synthesized particles: • The smallest particles are obtained with a moderate ΔT, while extreme ΔT values produce larger particles with greater degrees of deswelling.

25° C 45° C 55° C **1.02** 0.6° C $1.0 0.6^{\circ}$ C 0.6° C **1.00** 8.1°C 8.1° C 8.1°C 15.6° C 15.6° C **0.98** 0.8 **0.96** Probability
0.4
0.4 **0.94** $\mathbf{\Omega}$ **0.92 0.90 25^oC 0.88** $0.2₁$ **45^oC 0.86 55^oC** 0.0 **0.84** 300 600 900 600 900 300 600 900 300 **0 2 4 6 8 10 12 14 16 18 20 CONTIN size distribution (nm)** ^D**T (^oC)** • Carefully fitting the DLS correlation functions with spectral 2000 $\frac{\Delta T}{0.6}$ $\frac{25C}{9.5}$ $\frac{55C}{1}$

1750 $\frac{3.1}{5.1}$ $\frac{1}{8.1}$ $\frac{1}{14.6}$ $\frac{1}{9}$ line shape analysis⁴ yielded non-trivial dependences of **1750** the stretching parameter, β, on ΔT: **1500** o Low ΔT: Swollen particles are relatively polydisperse, while collapsed particles are more monodisperse \sim 1250 $\frac{1}{14}$ 6

 $ΔT (°C)$ o Low ΔT: Low density, polydisperse swollen particles that collapse to monodisperse, hard spheres above T_V

References and Acknowledgements

References:

1: Flory, P. J. **1953**. Principles of Polymer Chemistry. Cornell University Press, Ithaca.

2: Xia, Xiaohu, et al. "Formation and Volume Phase Transition of Hydroxypropyl Cellulose Microgels in Salt Solution." *Macromolecules*, vol. 36, no. 10, **2003**, pp. 3695–3698., doi:10.1021/ma0216728.

- We obtained the relative swollen particle densities with two methods:
	- Flory Huggins fits to DLS shrinkage curves yield a theoretical # of chains per particle, ν
	- SLS experiments, when performed at 25°C with unchanged optical settings across a series of samples, yield the relative molecular weights of the swollen particles
	- Dividing the SLS-obtained M_W with the particle volume (obtained by assuming a spherical shape and using the measured R_h values reported above) yields a fully-experimental swollen particle density

3: Hirotsu, Hirokawa, Tanaka, et al. "Volume‐Phase Transitions of Ionized N‐Isopropylacrylamide Gels." *J. Chem. Phys.*, vol. 87, no. 2, **1987**, pp. 1392–1395., doi:10.1063/1.453267.

4: Phillies, G. D. J.; O'Connell, R.; Whitford, P.; Streletzky, K.A. J. Chem. Phys. 119, 9903-9913 (2003); Streletzky, K. A.; McKenna, J. T.; Mohieddine, R. J. Polym. Sci. Part B-Polym. Phys. 2008, 46, 771-781.

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Modelling particle shrinkage with Flory Huggins

- o Intermediate ΔT: High density, monodisperse particles above and below T_{V}
- o High ΔT: Low density, polydisperse swollen particles that remain relatively polydisperse after collapse above T_V

3) Particle density

2) Particle polydispersity

1) Particle size and swelling/loading capacity

^D**T (^oC)**

 $\Delta T(^{\circ}C)$

• Dynamic light scattering yielded the hydrodynamic radius, R_h , for each of the samples synthesized at varying synthesis temperatures ($\Delta T = T_{\rm synthesis} - T_{\rm C}$). We show these results at three solution temperatures above.

Parent polymer and synthesized microgel particles

0 2 4 6 8 10 12 14 16 18 20 0 2 4 6 8 10 12 14 16 18 20

Flory Huggins Mean Field Theory¹ :

- $=\frac{1}{k_B\left[\frac{v_1v}{N_A\phi^2}\left[\left(\frac{\phi}{\phi_0}\right)-2\left(\frac{\phi}{\phi_0}\right)^{1/3}\right]-\frac{2}{\phi}-\frac{2\ln(1-\phi)}{\phi^2}\right]+\Delta S}$
- In agreement with the literature^{2,3}, we found that fixing one parameter in the Flory Huggins model produced better (more stable, more consistent) fits to the experimental data
- To study the effect of $\Delta T = T_{syn} T_C$ on particle density, v, we opted to fix either ΔH or ΔS
	- Since the ratio $\Delta H / \Delta S$ is a close approximation³ of the constant volume phase transition temperature, T_ν, the choice of fixing either ΔH or ΔS is inconsequential. We chose ΔS.
- **700 55 T**_V \approx dH/dS (Tanaka 1987) **(^oC)** • Fitting shrinkage curves with the Flory Huggins model yields three important results: • dH/dS approximation of T_V • Maximum slope measure of T_V • Particle density, ν
- o Intermediate ΔT: Particle polydispersity is unchanged between swollen and collapsed states
- o High ΔT: Both swollen and collapsed particles are relatively polydisperse, with unpredictable trends
- CONTIN distributions reveal similar polydispersity trends
- Multi-angle DLS suggests that while polydispersity varies, all particles are essentially spherical. The difference is likely due to differences in particle structure and density.

